21. Davies and Ludlam, "Origin of Laminated," p. 3527-3546.

22. Under some circumstances thin layers of calcium carbonate are formed as a result of the action of microscopic forms of algae. However, the amount of layer thickness which can be formed in one season in this manner is not essentially different from the amount which can precipitate inorganically under ideal conditions.

23. A. M. Klingspor, "Middle Devonian Muskeg Evaporites of Western Canada," American Association of Petroleum Geologists Bulletin, v. 53 (1969), p. 928-938.

24. The origin of salt domes (composed of common salt) by being pushed up out of the deeper layers of the earth's crust is entirely different from the production of a series of alternating salt, calcite, and organic layers in a marine basin. In the many ancient marine basins of the world which have such deposits, the fact that the common salt is in thin uniform layers, alternating with anhydrite--or with anhydrite and calcite or dolomite--is more than adequate proof that the layers originated by slow precipitation, rather than by being pushed up from beneath. And, because of the fact that fossiliferous carbonate strata underlie them, we know that they were formed after the creation of life on the earth.

25. Davies and Ludlam, "Origin of Laminated," p. 3528, 3532, 3534-3535.

26. Ibid., p. 3535-3537.

27. Davies does not state the total number of laminations and other layers in this deposit; but even the lower 80 feet, which contains approximately 40 feet of very thin laminations, must have required at least 10,000 years for its formation. This is allowing for normal rates of evaporation from a stagnant body of water, and taking into consideration the fact that the amounts of calcium carbonate and anhydrite which are present in the water awaiting precipitation are small. The amount of water which evaporates from ocean or lagoon surfaces annually in the most arid and hot regions of the world rarely exceeds 5 meters depth, and is usually only about one-half that amount. Even the 5 meters of water has only enough minerals to produce a layer of calcium carbonate 0.5 mm in thickness, and a 2.2 mm layer of anhydrite (W. B. F. Ryan, et al., Initial Reports of the Deep Sea Drilling Project, v. XIII, 1973, p. 1214). These quantities are for "normal" or average sea water. It can be seen from Table 1 (which is also based on normal sea water) that even if the water were completely saturated with respect to $CaCO_3$, only about 90% more calcium carbonate could be precipitated per unit of water. Thus in the above case the amount of calcium carbonate precipitated from 5 meters of water would still be less than 1 mm.

28. D. J. Kinsman, "Modes of Formation, Sedimentary Associations, and Diagnostic Features of Shallow-Water and Supratidal Evaporites," American Association of Petroleum Geologists Bulletin, v. 53 (1969), v. 53 (1969), p. 832, 839.